

AUSTRALIA

SPRUSON & FERGUSON

PATENTS ACT 1990

PATENT REQUEST: STANDARD PATENT

I/We, the Applicant(s)/Nominated Person(s) specified below, request I/We be granted a patent for the invention disclosed in the accompanying standard complete specification.

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[54] Invention Title:

Mass Catalysts Based on Chromium and Nickel Oxides and their
Application to the Fluorination of Halogenated Hydrocarbons

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SPRUSON & FERGUSON

Patents Act 1990

NOTICE OF ENTITLEMENT

I, John Gordon Hinde, of Spruson & Ferguson, St. Martins Tower, 31 Market Street, Sydney, New South Wales, 2000, Australia, being the patent attorney for the Applicant/Nominated Person in respect of Application No 29972/92 state the following:-

The Applicant/Nominated Person has entitlement from the actual inventor(s) as follows:-

The Applicant/Nominated Person, by virtue of a Contract of Employment between the actual inventor(s) as employee(s) and the Applicant/Nominated Person as employer, is a person which would be entitled to have the patent assigned to it if a patent were granted on an application made by the actual inventor(s).

The Applicant/Nominated Person is the applicant of the basic application(s) listed on the Patent Request. The Applicant/Nominated Person having made the basic application in its former name of Atochem, the name of the Company now having been changed.

The basic application(s) listed on the Patent Request is/are the application(s) first made in a Convention country in respect of the invention.

DATED this 18th day of January 1993


.....
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MASS CATALYSTS BASED ON CHROMIUM AND NICKEL OXIDES AND THEIR APPLICATION TO THE FLUORINATION OF HALOGENATED HYDROCARBONS
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- (57) Claim

1. A mass catalyst based on chromium and nickel oxides having a Ni/Cr atomic ratio of between 0.05 and 5 which catalyst is obtained by a process comprising:

a) forming a sol of chromium (III) and nickel (II) hydroxides,

b) gelling this sol, and

c) drying and calcining the product to a temperature of between 250 and 450°C.

15. A process for fluorination of a saturated or olefinic halogenated hydrocarbon which process comprises fluorinating the halogenated hydrocarbon by HF in the gas phase, using a catalyst as claimed in any preceding claim.

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COMPLETE SPECIFICATION

FOR A STANDARD PATENT

ORIGINAL

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Invention Title:

Mass Catalysts Based on Chromium and Nickel Oxides and
their Application to the Fluorination of Halogenated
Hydrocarbons

The following statement is a full description of this invention, including the
best method of performing it known to me/us:-

The present invention relates to the fluorination of the halogenated hydrocarbons by gas phase catalysis and more particularly relates to new mass catalysts based on chromium and nickel and their application to the synthesis of hydrohaloalkanes.

The intensive research currently conducted on substitutes for chlorofluorocarbons (CFCs) is directed, inter alia, towards the synthesis of hydrohaloalkanes. Some steps of this synthesis may be carried out by fluorination with hydrofluoric acid by heterogeneous gas phase catalysis.

Numerous metal compounds (for example chromium, cobalt, nickel, iron, copper, manganese, and the like) have a catalytic effect for these fluorination reactions. The catalysts proposed in the literature are either mass catalysts or supported catalysts, the support being generally carbon or alumina (partially converted to AlF_3 after fluorination).

Amongst this second category of catalyst, a large variety of metal compounds are found and there are numerous patents describing fluorination processes based on this type of catalyst. Thus, mention may be made of US patents 2 744 147 and 2 744 148, which describe the fluorination of a haloalkane on a catalyst based on a metal (chromium, cobalt or nickel for US patent 2 744 147 and chromium, cobalt, nickel, copper or palladium for US patent 2 744 148) supported on alumina.

More recently, European patent EP 0 366 797

describes a fluorination process using a catalyst consisting of at least one metal (nickel, cobalt, iron, manganese, chromium, copper and silver) fluoride supported on an alumina having a significant mesoporosity.

5 For all of these catalysts, the support imparts a certain solidity. However, as the amount of active material is smaller than in a mass catalyst, the catalytic activity may be affected. Moreover, the low contents of non-precious metals do not enable their recovery in an economic manner
10 from spent catalysts to be envisaged.

Mass fluorination catalysts are generally based on chromium and the starting materials used for their development are very varied (salts, oxides, halides and the like).

15 Thus, US patent 4 912 270 and European EP 0 313 061 claim fluorination processes using catalysts based on chromium oxide which are obtained, respectively, by reduction of chromium trioxide by means of an alcohol and by pyrolysis of ammonium bichromate.

20 French patent FR 2 135 473 describes the preparation of a catalyst containing chromium and nickel and its use in the fluorination of functional perhalogenated compounds. This catalyst is obtained by thermal decomposition of organic chromium and nickel salts. The
25 nickel contents remain low since the Ni/Ni + Cr atomic ratio is still less than 0.1.

Japanese patent publications JP 2-172932/90 and

2-172933/90 describe, respectively, the fluorination of 1,1-difluoro-1,2,2-trichloroethane (F122) and of 1-chloro-2,2,2-trifluoroethane (F133a) on a chromium catalyst to which a doping metal has been added which enables the reaction

5 temperature to be lowered while maintaining a significant activity and therefore improves the life of the catalyst by restricting the crystallisation of the chromium. The use of chromium/nickel combinations is not illustrated by the examples in these publications.

10 French patent FR 2 501 062 describes the preparation of a mass chromium oxide in the form of microspheres having a diameter of between 0.1 and 3 mm. This catalyst is obtained by gelling a chromium hydroxide sol in a solvent which is immiscible or partially miscible with water.

15 The product obtained is highly solid and is particularly suitable for fluorination reactions in a fluidised bed.

The disadvantage of catalysts based on chromium oxide is their low resistance at high temperature (350-500°C) to the crystallisation which contributes to reducing their
20 life.

Moreover, these chromium-based catalysts promote the oxidation of hydrochloric acid by the oxygen dissolved in the reagents or deliberately introduced. Water and chlorine form by a Deacon reaction (Chemical Week 1987, 24 June, page
25 18) and these in turn react with the organic compounds, which leads to the formation of non-valorisable by-products and, consequently, to a lowering of the selectivities.

It has now been found that the addition of a nickel compound to a chromium derivative in order to form a sol of chromium and nickel hydroxides makes it possible, while retaining the advantages of a mass catalyst, not only to prolong the life of the catalyst by improving the resistance of the chromium-based compound to crystallisation but also to improve the selectivities in gas phase fluorination reactions owing to a partial inhibition of the oxidation of hydrochloric acid in the presence of chromium.

10 According to the present invention there is provided a mass catalyst based on chromium and nickel oxides, obtained by a process comprising:

- a) forming a sol of chromium (III) and nickel (II) hydroxides,
- 15 b) gelling this sol, and
- c) drying and calcining the product to a temperature of between 250 and 450°C.

The invention also provides a process for the fluorination of saturated or olefinic halogenated hydrocarbons by HF in the gas phase using such a catalyst.

20 In the catalysts according to the invention, which may be in various forms (spheres, extruded products, pellets, and the like) the Ni/Cr atomic ratio may range from 0.05 to 5. It is advantageously between 0.1 and 3.5 and preferably between 0.15 and 3.

Depending on the gelling technique used (in form of droplets or in mass), the catalyst precursor, consisting

of a homogeneous mixture of chromium and nickel hydroxides, is obtained, which, after drying, is either in the form of microspheres or in the form of a powder which may be shaped using well-known techniques, for example by extrusion or by
5 pelleting. After calcination, a mass catalyst consisting of a homogeneous mixture of chromium and nickel oxides is obtained.

The sol of chromium (III) and nickel (II) hydroxides may be formed in a manner known per se from
10 chromium and nickel precursors.

Chromium precursors which may be mentioned are chromium oxides, hydroxides, halides, oxyhalides, acetates, nitrates and sulphates, but it is also possible to use any other chromium compound capable of leading to a chromium
15 hydroxide sol. The preferred precursors are chromium salts such as the sulphates, chlorides, acetates and nitrates, with chromium (III) sulphate or acetate being particularly preferred.

Nickel precursors which may be mentioned are the
20 hydroxides, oxyhalides, nitrate, acetate and sulphates of this metal, but it is also possible to use any other nickel compound which is soluble in water and capable of forming a sol or of being incorporated in the chromium gel. The preferred precursors are the highly soluble salts, such as
25 nickel nitrates and especially nickel chlorides or sulphates.

With certain precursors of chromium and nickel, particularly the oxides, hydroxides, acetates and sulphates,

the sol may be formed at room temperature. However, where the precursor is a chromium nitrate or halide, the formation of the sol needs a heating step at a temperature between 60 and 100°C, preferably between 80 and 95°C. Furthermore, even 5 when the nature of the precursor permits the sol to be formed at low temperature, it is advantageous to heat the solution of precursors at a temperature of between 60 and 100°C, preferably between 80 and 95°C.

The formation of the sol may also be improved by 10 adding to the aqueous solution of the precursors a complexing agent for chromium and/or nickel such as, for example, ammonium acetate, sulphate or phosphate, in a molar amount which may be up to five times the total number of moles of the chromium and nickel precursors.

15 Various additives may be added to the sol in order to improve the physicochemical and catalytic properties of the final catalyst. Thus, it is possible to add (% relative to the weight of the sol):

a) 2 to 10 % of Cr_2O_3 or $\text{Cr}_2\text{O}_3 \cdot 2\text{H}_2\text{O}$ powder 20 previously dried at 300°C in order to increase the mechanical strength of the microspheres;

b) 0.1 to 3 % of alumina monohydrate in order to increase the resistance to attrition of the catalyst;

c) 5 to 25 % of hexamethylenetetramine and/or 25 urea, the presence of which leads, by decomposition at the gel temperature of the sol, to an additional evolution of ammonia;

d) 1 to 10 % of colloidal silica in order to increase the porosity of the catalyst (the silica is removed in the form of silicon tetrafluoride during the treatment with hydrofluoric acid);

5 e) other additives, such as wetting agents (for example lauryldiethanolamide or polyethylene glycol monostearate) or thickeners (for example hydroxymethylcellulose or microcrystalline cellulose) in order to improve the spherical nature of the microspheres.

10 Using the sol of chromium and nickel hydroxides as the starting material, the mixed chromium and nickel oxide catalysts according to the invention may be obtained in the form of microspheres by gelling the sol in the following way:

1) dispersing the sol in the form of droplets in
15 an organic solvent which is immiscible or sparingly miscible with water and gelling at elevated temperature;

2) receiving the microspheres formed in the gelling solvent or in an ammonia solution in order to complete gelling:

20 3) washing the microspheres with dilute ammonia and with water in order to remove the impurities, and optionally drying.

First step

The formation and the gelling of the microspheres
25 may be carried out in a column having a height of between 1 and 6 m, which column is extended by a release zone and has a stream of hot solvent passing through it from bottom to top.

The sol of chromium and nickel hydroxides is injected at the top of the column by means of a tube of small diameter arranged concentrically inside another tube of larger diameter, through which the solvent cooled to a temperature not exceeding 30°C (preferably 25°C) arrives in co-current in order to prevent gelling in the injectors.

The diameter of the injector and the flow of the sol determine the dispersion of the droplets and consequently the final size of the microspheres. During its injection, it is expedient to keep the sol at a temperature of between 2 and 8°C, preferably between 3 and 5°C, so as to prevent variations in its viscosity. These variations are reflected in a change in the sol and, consequently, in the physicochemical properties of the microspheres.

It is advantageous to partially neutralise the sol, before its injection, with an amount of ammonium hydroxide sufficient to obtain a viscosity adapted to the dispersion of the droplets. This amount depends on the nature of the chromium and nickel precursors used. The $\text{NH}_4\text{OH}/\text{Cr}+\text{Ni}$ molar ratio is advantageously less than 2 and preferably less than 1.5.

The ascending stream of hot solvent (temperature of between 25 and 140°C) is introduced at the base of the column at an ascending speed of between 0.1 and 5 m/s. This speed determines the residence time of the microspheres and, consequently, their degree of gelling.

In its upper part, the column may be provided with

a second inlet for solvent cooled to a temperature not exceeding 30°C. This facilitates the maintenance of a temperature gradient in the column and enables the gelling rate to be controlled.

5 The organic solvent which is immiscible or sparingly miscible with water may be chosen from alcohols, such as butanol, hexanol, 2-ethylpentanol and 2-ethylhexanol, this latter alcohol being particularly preferred.

Second step

10 On leaving the bottom of the column, the microspheres are collected in a vessel containing the organic gelling solvent or, preferably, containing aqueous ammonia, the concentration of which may be between 0.1 and 14 N and the temperature of which is kept between 15 and 70°C,
15 preferably between 25 and 50°C.

... The presence of ammonia in the vessel effectively
... completes the gelling of the microspheres, thus increasing
... their mechanical strength and also ensuring the first
... washing. An increase in the concentration of the ammonia
20 solution enables the solidity and the density of the
microspheres to be increased.

Third step

... The microspheres are suitably washed several times
... with dilute ammonia (0.1 to 1 N) and then with water, before
25 optionally being dried in air at a temperature of between 80 and 200°C. However, this drying at low temperature may be omitted without risk of altering the properties of the

microspheres.

The microspheres are then calcined under inert gas (for example nitrogen, argon or the like) or in air at a temperature of between 250 and 450°C, preferably between 300 5 and 420°C.

Instead of being gelled in the form of microspheres in accordance with the method described above, the sol of chromium and nickel hydroxides may be completely gelled by adding a base, preferably aqueous ammonia, until 10 the sol of hydroxides solidify. The product obtained is then suitably washed with dilute ammonia (0.1 to 1 N) and/or with water, before being dried at a temperature of between 80 and 200°C. The washings and filtrations may be made easier by adding from 5 to 100 ppm of a flocculant, preferably a 15 polyacrylamide. The catalyst is then calcined in the same manner as described above, this calcination being able to be preceded or followed by a shaping step using known techniques (extrusion, pelleting).

The mass catalysts according to the invention may 20 be used for the fluorination of saturated or olefinic halogenated hydrocarbons by HF in the gas phase. They are particularly suitable for the fluorination of halogenated hydrocarbons leading to fluorinated C₁ to C₃ compounds containing one or more hydrogen atoms. The following 25 compounds may be mentioned, without any limitation being implied, as examples of starting halogenated hydrocarbons: CHCl₃, CCl₂ = CHCl, CHCl₂-CClF₂, CH₂Cl-CF₃, CH₃-CCl₂-CH₃,

$\text{CCl}_3\text{-CF}_2\text{-CH}_3$, $\text{CCl}_3\text{-CF}_2\text{-CHCl}_2$, $\text{CCl}_3\text{-CF}_2\text{-CH}_2\text{Cl}$, $\text{CHCl}_2\text{-CHCl-CH}_3$, and also $\text{CCl}_2 = \text{CCl}_2$; this latter compound does not contain hydrogen, but the addition of HF leads to hydrohalogenated compounds.

5 In order to work at the optimum activity, the catalyst requires a treatment with hydrofluoric acid, which is undiluted or diluted with nitrogen. Although the presence of nickel slows down the crystallisation of the catalyst, an activation of this type may locally generate temperatures
10 higher than 500°C . It is for this reason that it is recommended to control the exothermicity of the activation by adjusting the addition of a diluent for HF and by starting this treatment at low temperature ($150\text{-}250^\circ\text{C}$). On the other hand, after passing through "exothermicity waves" in the
15 catalyst bed, it is advised that the temperature be progressively increased in order to reach a maximum of $350\text{-}450^\circ\text{C}$ at the end of activation.

The fluorination reaction temperature depends on the reaction studied and, of course, on the desired reaction
20 products. Thus, for a partial replacement of chlorine atoms by fluorine, the reaction is carried out at temperatures of between 50 and 350°C ; the replacement of all of the chlorine atoms may require temperatures of between 300 and 500°C .

The contact time also depends on the reaction
25 studied and the desired products. In general it is between 3 and 100 seconds; however, in order to obtain a good compromise between high degree of conversion and high

productivity, this contact time generally remains less than 30 seconds.

The HF/organic compound molar ratio is also linked to the reaction studied: it depends, inter alia, on the stoichiometry of the reaction. In the majority of cases it may vary between 1/1 and 20/1, but, in this case also, in order to obtain high productivities it is often less than 10.

The operating pressure is preferably between 1 and 20 bars absolute (0.1 to 2 MPa).

10 The catalysts according to the invention may operate in a fixed bed or in a fluidised bed. The catalysts in the form of microspheres are very solid and therefore particularly suitable for reactions in a fluidised bed.

15 Catalysts for which the activity has fallen as a consequence of contamination may be regenerated by cleaning the catalyst with a compound capable of oxidising and converting the products (organic products, coke, and the like) deposited on the catalyst into volatile products. In this capacity, oxygen or a mixture containing oxygen (air for 20 example) is perfectly suitable and enables the initial catalyst activity to be restored.

25 In order to ensure the regeneration of the catalyst without inducing crystallisation it is recommended to carry out this treatment at a temperature of between 200 and 450°C and more particularly between 300 and 380°C. Similarly, it is expedient to limit the exothermicity of this "combustion" by controlling the oxygen flow rate (at the

start of regeneration, low flow rate of oxygen diluted in an inert gas) so as to maintain a temperature of below 450°C.

In order to maintain the activity of the catalyst, it is also possible to carry out the fluorination reaction in the presence of oxygen introduced in an O₂/organic compound molar ratio which can range from 0.001 to 0.05 and is preferably between 0.005 and 0.03.

The following Examples illustrate the invention without restricting it. The pore volume in Examples 1 to 12 was determined by mercury porosimetry and corresponds to that of the pores having a radius between 4 nm and 63 μm.

PREPARATION OF THE CATALYSTS

EXAMPLE 1 Catalyst A

The procedure is carried out in the apparatus 15 which is shown in appended Figure 1 and comprises the following main elements:

- a glass reactor (1) which has a volume of 3 litres, for the preparation of the sol of chromium and nickel hydroxides; it is provided with a double wall and fitted with a biconical stirrer (11) rotating at 3000 rev/min; a pump and a heat exchanger (12) make it possible to remove the calories from the neutralisation and thus to keep the sol at the desired temperature;
- a glass column (2) for gelling the sol in the form of microspheres; this column (80 mm in diameter and 1.5 m high) is extended at the top by a release zone (21) 100 mm in diameter and 100 mm high;

- an injector device (3) consisting of a tube having an internal diameter of 2 mm for the injection of the sol into the column 2; this tube is arranged concentrically inside a tube having an internal diameter of 12 mm, through which 2-ethylhexanol is injected;

- a receiving tank (4) for collecting the microspheres formed in the column 2; this tank, which has a volume of 5 litres, contains aqueous ammonia which is homogenised and kept at the desired temperature by means of a circuit comprising a pump and a heat exchanger (41);

- a reservoir (5) for feeding the injector device 3 and the column 2 with 2-ethylhexanol serving as dehydrating solvent.

a) Preparation of the sol of chromium and nickel hydroxides

15

An aqueous chromium and nickel solution is prepared at room temperature by dissolving 550 g of basic chromium sulphate $\text{Cr}_2(\text{SO}_4)_2(\text{OH})_2\text{Na}_2\text{SO}_4$ and 238 g of nickel chloride hexahydrate $\text{NiCl}_2 \cdot 6\text{H}_2\text{O}$ in 367 g of water. 70 g of crystalline chromium (III) oxide are then added to this aqueous solution.

The mixture is cooled to 5°C and, while maintaining this temperature, the following are added successively:

25

- 151.5 g of cold 11 N ammonia (5°C),
- a cold (5°C) aqueous solution comprising 204 g of hexamethylenetetramine, 10.8 g of urea and 300 ml of

water, and

- 225 g of a 30 % silica sol in water (CECASOL from CECA).

b) Injection of the sol and synthesis of the microspheres

5 The sol of chromium and nickel hydroxides, prepared in step a), is then routed towards column 2 through the line (13) by means of a pump (14). The use of a peristaltic pump enables any variation in flow rate to be
10 avoided. For significant lengths of routing line 13, it is necessary to sheath the line in order to prevent refrigerating losses.

The sol at 5°C is injected at a rate of 0.6 l/h through the tube having an internal diameter of 2 mm, while
15 2-ethylhexanol at a temperature of 25°C is injected through the line 51 at a flow rate of 10 l/h by means of the tube having an internal diameter of 12 mm in order to prevent premature gelling of the sol in the injector device 3; the end of the latter is immersed about 5 mm into the organic
20 solvent. A flow of 2-ethylhexanol at the same temperature (25°C) may be passed into the upper part of the column through a line 52, the inlet of which is located below the release zone.

Another stream of 2-ethylhexanol, heated to a
25 temperature of 120°C, is introduced at a flow rate of 45 l/h into the bottom of the column through the line 53. This stream of 2-ethylhexanol passes through the column from

bottom to top at an ascending rate of about 9 m/h and exits via an overflow located at the top of the release zone. The two 2-ethylhexanol flows enable a temperature gradient, which is necessary to control the gelling and dehydration kinetics of the sol, to be obtained in the column.

The 2-ethylhexanol leaving the column is passed to the reservoir 5 through the line 54. In a device which is not shown in Figure 1, it is pre-purified by washing with water and distillation in order to remove the dissolved ammonium sulphate, the organic residues (urea, hexamethylenetetramine and formol) and the water extracted from the sol.

At the lower end of the column, the microspheres are collected in the receiving tank 4 containing 1 N ammonia homogenised and kept at 40°C by means of the closed circuit 41.

The production is about 100 g/h per injector, but may be increased by using several injectors.

The microspheres are washed thoroughly with dilute ammonia (0.1 N) and then with water and dried at 120°C. They are then calcined at 420°C under a nitrogen atmosphere for 4 hours.

The catalyst A thus obtained has the following characteristics:

25

- Ni/Cr atomic ratio = 0.34
- diameter of the microspheres = 0.5 to 2.5 mm
- apparent density = 1.25 g/cm³

- specific surface area (BET) = $185 \text{ m}^2/\text{g}$
- pore volume = $0.05 \text{ cm}^3/\text{g}$
- bulk crush strength (BCS) = $2.1 \text{ kg}/\text{cm}^2$

EXAMPLE 2 Catalyst B

5 The procedure is as in Example 1, except that the ammonia solution in the receiving tank for the microspheres has a concentration of 11 N instead of 1 N.

 The characteristics of the catalyst B thus obtained are as follows:

- 10 - Ni/Cr atomic ratio = 0.27
- diameter of the microspheres = 0.5 to 2.5 mm
- apparent density = $1.4 \text{ g}/\text{cm}^3$
- specific surface area (BET) = $102 \text{ m}^2/\text{g}$
- pore volume = $0.07 \text{ cm}^3/\text{g}$
- 15 - bulk crush strength (BCS) = $7.5 \text{ kg}/\text{cm}^2$

EXAMPLE 3 Catalyst C

 The procedure is as in Example 1, but without addition of silica sol; a catalyst C having the following characteristics is obtained:

- 20 - Ni/Cr atomic ratio = 0.25
- diameter of the microspheres = 0.5 to 2.5 mm
- apparent density = $1.5 \text{ g}/\text{cm}^3$
- specific surface area (BET) = $25 \text{ m}^2/\text{g}$
- pore volume = $0.09 \text{ cm}^3/\text{g}$

EXAMPLE 4 Catalyst D

The procedure is as in Example 1, reducing the nickel content, that is to say using an aqueous solution prepared by dissolving 550 g of basic chromium sulphate and 5 45 g of nickel chloride hexahydrate in 440 g of water as the starting material.

The catalyst D thus obtained has the following characteristics:

- Ni/Cr atomic ratio = 0.07
- diameter of the microspheres = 0.5 to 2.5 mm
- apparent density = 1.38 g/cm³
- specific surface area (BET) = 187 m²/g
- pore volume = 0.044 cm³/g

EXAMPLE 5 Catalyst E

The procedure is as in Example 1, except that the microspheres are calcined at 350°C instead of 420°C.

The catalyst E thus obtained has the following characteristics:

- Ni/Cr atomic ratio = 0.37
- diameter of the microspheres = 0.5 to 2.5 mm
- apparent density = 1.1 g/cm³
- specific surface area (BET) = 178 m²/g
- pore volume = 0.09 cm³/g

EXAMPLE 6 Catalyst F (Comparative)

The procedure is as in Example 1, omitting the

nickel. The starting aqueous solution consists solely of basic chromium sulphate (550 g) and water (475 g).

The characteristics of this catalyst F, prepared for comparison, are as follows:

- 5 - diameter of the microspheres = 0.5 to 2.5 mm
- apparent density = 1.04 g/cm³
- specific surface area (BET) = 209 m²/g
- pore volume = 0.052 cm³/g
- bulk crush strength (BCS) = 3.8 kg/cm²

10 If the ammonia contained in the receiving tank is replaced by 2-ethylhexanol at 120°C by modifying the apparatus in accordance with appended Figure 2, a product is obtained which has the following characteristics:

- apparent density: 1.07 g/cm³
- 15 - specific surface area: 203 m²/g
- bulk crush strength (BCS): 2.3 kg/cm²

EXAMPLE 7 Catalyst G (Comparative).

The procedure is as in Example 6 except that the microspheres are calcined at 350°C instead of 420°C.

20 The catalyst thus obtained has the following characteristics:

- diameter of the microspheres = 0.5 to 2.5 mm
- apparent density = 1.3 g/cm³
- specific surface area (BET) = 181 m²/g
- 25 - pore volume = 0.05 cm³/g

EXAMPLE 8 Catalyst H (Comparative)

The procedure is as in Example 1 without the addition of either nickel or silica sol. The starting aqueous solution consists solely of basic chromium sulphate 5 (550 g) and water (475 g).

The characteristics of this catalyst H, prepared for comparison, are as follows:

- diameter of the microspheres = 0.5 to 2.5 mm
- apparent density = 1.4 g/cm³
- 10 - specific surface area (BET) = 10 m²/g
- pore volume = 0.28 cm³/g

EXAMPLE 9 Catalyst I

A solution of 200 g of chromium nitrate nonahydrate and 118.8 g of nickel chloride hexahydrate in 15 1000 g of water is prepared and heated at 80°C for 2 hours. After cooling to 20°C, this solution is then gelled by adding 190 ml of 14.7N aqueous ammonia.

The gel obtained is then washed twice with 450 ml of 0.1N aqueous ammonia, then twice with 450 ml of distilled 20 water, each washing step being followed by a filtration.

The powder obtained is then dried for 14 hours at 100°C under vacuum (20 kPa) and then calcined at 350°C under a nitrogen atmosphere for 4 hours.

The characteristics of the catalyst I thus 25 obtained are as follows:

- specific surface area (BET) = 160 m²/g

- pore volume = $0.74 \text{ cm}^3/\text{g}$

EXAMPLE 10 Catalyst J (comparative)

The procedure is as in Example 9 except that the solution of chromium and nickel salts is heated at 40°C instead of 80°C .

The catalyst J thus obtained has the following characteristics:

- specific surface area (BET) = $139 \text{ m}^2/\text{g}$
- pore volume = $0.60 \text{ cm}^3/\text{g}$

10

EXAMPLE 11 Catalyst K

An aqueous chromium and nickel solution is prepared by dissolving 274 g of nickel chloride hexahydrate in 200 g of water and then by adding 275 g of basic chromium sulphate and 183.5 g of water.

15

112 g of CECASOL are added to this solution and 390 ml of an 11 N ammonia solution are then added until the sol sets completely. The gel obtained is washed with a dilute ammonia solution and with water before being dried at 120°C for 15 hours and calcined at 250°C for 4 hours.

20

A 12 % aqueous polyvinyl alcohol solution is added to the powder collected and the mixture is then dried at 85°C for 15 hours. The mixture is then shaped by pelleting and calcined at 420°C for 4 hours.

The catalyst K thus obtained has the following characteristics:

Cr = 24.7 % by mass

Ni = 20.15 % by mass

Ni/Cr atomic ratio = 0.42

specific surface area (BET) = 54.8 m²/g

5 pore volume = 0.162 cm³/g

EXAMPLE 12 Catalyst L (Comparative)

A chromium and nickel solution is prepared by mixing at room temperature 80 ml of a 1 M Cr(NO₃)₃·9H₂O solution and 80 ml of a 1M NiCl₂·6H₂O solution. The mixture
10 is then neutralised with 50 ml of 14 N ammonia and the product, recovered by centrifuging, is dried at 100°C for 15 hours.

0.4 g of graphite and 5.1 g of a 12 % aqueous polyvinyl alcohol solution are added to the powder obtained
15 and the mixture is then dried at 100°C (15 hours) and shaped into pellets. The final product is calcined at 350°C for 4 hours.

The catalyst L thus obtained has the following characteristics:

20 Ni/Cr atomic ratio = 0.95

specific surface area (BET) = 82.8m²/g

pore volume = 0.17 cm³/g

FLUORINATION EXAMPLES

In the following examples:

25 - the percentages indicated are molar

- the hydrofluoric acid used is a commercial product containing only traces of water;

- the starting 1-chloro-2,2,2-trifluoroethane (F133a) is a 99.9 % pure product;

5 - the reactor used is a 250 ml Inconel tube heated by means of a fluidised bath of alumina.

The activation of the catalyst by means of HF is carried out in this reactor on a 100 ml sample. After drying for 3 hours under nitrogen (5 l/h) at 250°C, hydrofluoric
10 acid is added progressively, at the same temperature, to nitrogen for a period of 5 hours (2 mols of HF introduced in the course of five hours). After passing through exothermicity peaks, the HF flow rate is increased in order to reach 1 mol/h and the temperature is then brought to
15 350°C. A temperature plateau is observed under these conditions for 8 hours.

Before they are introduced into the reactor, the reagents are mixed and heated to the reaction temperature in an Inconel preheater.

20 After washing with water - in order to remove the hydracids - and drying over CaCl_2 , the reaction products are analysed on-line by gas phase chromatography.

EXAMPLES 13 TO 15

F133a fluorination tests are carried out under
25 atmospheric pressure, in the absence of oxygen, using catalysts A, D and F, activated by the procedure described

above. The fluorination results are summarised in Table 1.

It is found that the activity of catalyst F (chromium on its own) falls much more rapidly than that of catalysts A and D (nickel/chromium) and that the 5 selectivities are also less good.

TABLE 1

CATALYST	A		D		F comparative	
	Example 13		Example 14		Example 15	
10 Age of the catalyst (hours)	26	221	22	220	24	221
15 HF/F133a molar ratio	4.1	4.2	3.8	3.9	4.0	4.0
Contact time (s)	3.9	3.8	4.1	4.0	3.9	3.9
Temp. (°C)	350	350	350	350	350	350
20 Degree of conversion of F133a (%)	22.4	22.6	20.2	18.2	19.5	9.1
25 Selectivity for 134a (CF ₃ CH ₂ F) (%)	97.3	97.3	97.9	98.3	96.9	94.5
30 Selectivity for F1122 (CF ₂ =CHCl) (%)	1.3	1.3	1.0	1.1	1.5	3.3
35 Selectivity for F123 (CF ₃ CHCl ₂) (%)	0	0	0	0	0.3	1.1

TABLE 1 (continuation)

CATALYST	A		D		F comparative	
	Example 13		Example 14		Example 15	
5 Selectivity for F124 (CF_3CHClF) (%)	0.3	0.3	0.2	0.3	0.5	0.5
10 Selectivity for F125 (CF_3CHF_2) (%)	0.1	0	0	0	0	0
20 Selectivity for F143a (CF_3CH_3) (%)	0.3	0.3	0.2	0.3	0.3	0

EXAMPLES 16 AND 17

25

F133a fluorination tests are carried out under an absolute pressure of 1.5 MPa in the presence of oxygen by injection of air and with catalysts E and G activated in accordance with the procedure described above. The fluorination results are summarised in Table 2.

TABLE 2

CATALYST	E		G comparative	
	Example 16		Example 17	
5 Age of the catalyst (hrs.)	22	268	22	265
HF/F133a molar ratio	5.1	5.1	5.1	5.1
10 O ₂ /F133a molar ratio	0.01	0.01	0.01	0.02
Contact time (s)	19.6	19.6	19.6	19.7
Temperature (°C)	350	350	350	350
15 Degree of conversion of F133a (%)	23.2	24.7	21.6	19.7
Selectivity for F134a (%)	96.5	95.9	93.3	89.8
Selectivity for F1122 (%)	0.05	0.04	0.1	0.1
20 Selectivity for F123 (%)	0.5	0.4	2.1	3
Selectivity for F124 (%)	0.5	0.9	1.1	1.3
25 Selectivity for F125 (%)	0.2	0.4	0.3	0.2
Selectivity for F143a (%)	0	0.02	0.1	0.1

Entirely as before, catalyst E (Ni/Cr) (Example 16) gives rise to a more stable activity and to a better selectivity than catalyst G without nickel (Example 17).

EXAMPLES 18, 19 AND 20

These examples, which are summarised in Table 3 below, were carried out using catalysts C and H without silica and permit comparison of the results of F133a 5 fluorination under atmospheric pressure in the absence of oxygen.

After 548 hours in use, catalyst C of Example 19 was subjected to regeneration by treatment under air (1 mol/hour) at 300°C of 24 hours. Catalyst C regenerated in 10 this way was then used for the tests of Example 20.

Examination of the results enables the effect of nickel on the activity of the catalyst, despite the absence of silica, to be estimated.

It is also found that the regenerated catalyst C 15 (Example 20) has an activity comparable to that of catalyst C at the start of the test.

TABLE 3.

CATALYST	H comparative Example 18	C				Regenerated C	
		Example 19				Example 20	
Age of the catalyst (hrs.)	47	65	333	548	48(*)	245(*)	
HF/F133a molar ratio	3.9	4	4	4	4	4	
Contact time (s)	4	4.1	4.1	4.2	4.8	4.6	
Temperature (°C)	350	350	350	350	350	350	
Degree of conversion of F133a (%)	4	20.4	18.6	14.8	20.4	20.8	
Selectivity for F134a (%)	87.5	97.1	97.3	98	98.5	98.1	
Selectivity for F1122 (%)	5	1	1.1	1.4	1	1	
Selectivity for F123 (%)	2.5	0.5	0.3	0.3	0	0	
Selectivity for F124 (%)	1.3	0.5	0.3	0.3	0.2	0.2	
Selectivity for F125 (%)	0	0	0	0	0	0	
Selectivity for F143a (%)	1.3	0	0	0	0	0	

(*) after regeneration

EXAMPLE 21

This example, summarised in Table 4 below, was carried out at atmospheric pressure in the absence of oxygen, using the catalysts I and J prepared from sols heated at 80 5 and 40°C respectively.

TABLE 4

CATALYST	I	J (comparative)
Age of the catalyst (hrs.)	28	21
10 HF/133a molar ratio	4.4	4.3
Contact time (s)	3.8	3.9
Temperature (°C)	350	350
Degree of conversion of F133a (%)	21.2	16.3
15 Selectivity for F134a (%)	98.3	97.5
Selectivity for F1122 (%)	0.9	1.5
Selectivity for F123 (%)	0	0.1
Selectivity for F124 (%)	0.2	0.2
Selectivity for F125 (%)	0.2	0.1
20 Selectivity for F143a (%)	0.3	0.6

Examination of the results enables the effect of the temperature of preparation of the sol of chromium and nickel on the activity of the catalyst, to be estimated.

The claims defining the invention are as follows:

1. A mass catalyst based on chromium and nickel oxides having a Ni/Cr atomic ratio of between 0.05 and 5 which catalyst is obtained by a process comprising:

- a) forming a sol of chromium (III) and nickel (II) hydroxides,
- b) gelling this sol, and
- c) drying and calcining the product to a temperature of between 250 and 450°C.

2. A catalyst according to claim 1, in which the Ni/Cr atomic ratio is between 0.1 and 3.5.

3. A catalyst according to claim 1, in which the Ni/Cr atomic ratio is between 0.15 and 3.

4. A catalyst according to any preceding claim in which the sol of chromium and nickel hydroxides is obtained from chromium (III) sulphate, acetate or nitrate and nickel chloride, nitrate or sulphate.

5. A catalyst according to any preceding claim, in which the sol also contains at least one additive chosen from the following:

- crystalline Cr_2O_3 powder,
- alumina monohydrate,
- hexamethylenetetramine and/or urea,
- colloidal silica,
- wetting agents, and
- thickeners.

6. A catalyst according to any preceding claim in which the formation of the sol comprises a heating step at a temperature of between 60 and 100°C.

7. A catalyst according to any of claims 1 to 5, in which the formation of the sol comprises a heating step at a temperature of between 80 and 95°C.

8. A catalyst according to any preceding claim, in which the sol is formed in the presence of a complexing agent for chromium and/or nickel.

10 9. A catalyst according to any preceding claim, in which the sol is gelled by means of aqueous ammonia.

10. A catalyst according to any preceding claim, which is in form of microspheres, and in which the gelled sol of chromium and nickel hydroxides is obtained by:

15 (a) dispersing the sol in the form of droplets in an organic solvent which is immiscible or sparingly miscible with water, and gelling at elevated temperature;

(b) receiving the microspheres formed in the same organic solvent or in an ammonia solution; and

20 (c) washing the microspheres with dilute ammonia and then with water and optionally drying.

11. A catalyst according to claim 10, in which the organic solvent used is 2-ethylhexanol.

12. A catalyst according to any one of claims 1 to 9, which is in the form of pellets or extruded product.

13. A catalyst according to any preceding claim, in which the calcination is carried out at a temperature of

between 300 and 420°C.

14. A catalyst according to claim 1 substantially as described in any one of Examples 1 to 5 and 9 to 11.

5 15. A process for fluorination of a saturated or olefinic halogenated hydrocarbon which process comprises fluorinating the halogenated hydrocarbon by HF in the gas phase, using a catalyst as claimed in any preceding claim.

10 16. A process according to claim 15, in which the halogenated hydrocarbon is 1-chloro-2,2,2-trifluoroethane.

15 17. A process for fluorination of a saturated olefinic halogenated hydrocarbon according to claim 15 substantially as described in any of Examples 13, 14, 16, 19, 20 and 21.

18. A fluorinated hydrocarbon obtained by the process claimed in any of claims 15 to 17.

DATED this EIGHTH day of DECEMBER 1992

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29972/92

Fig. 1.

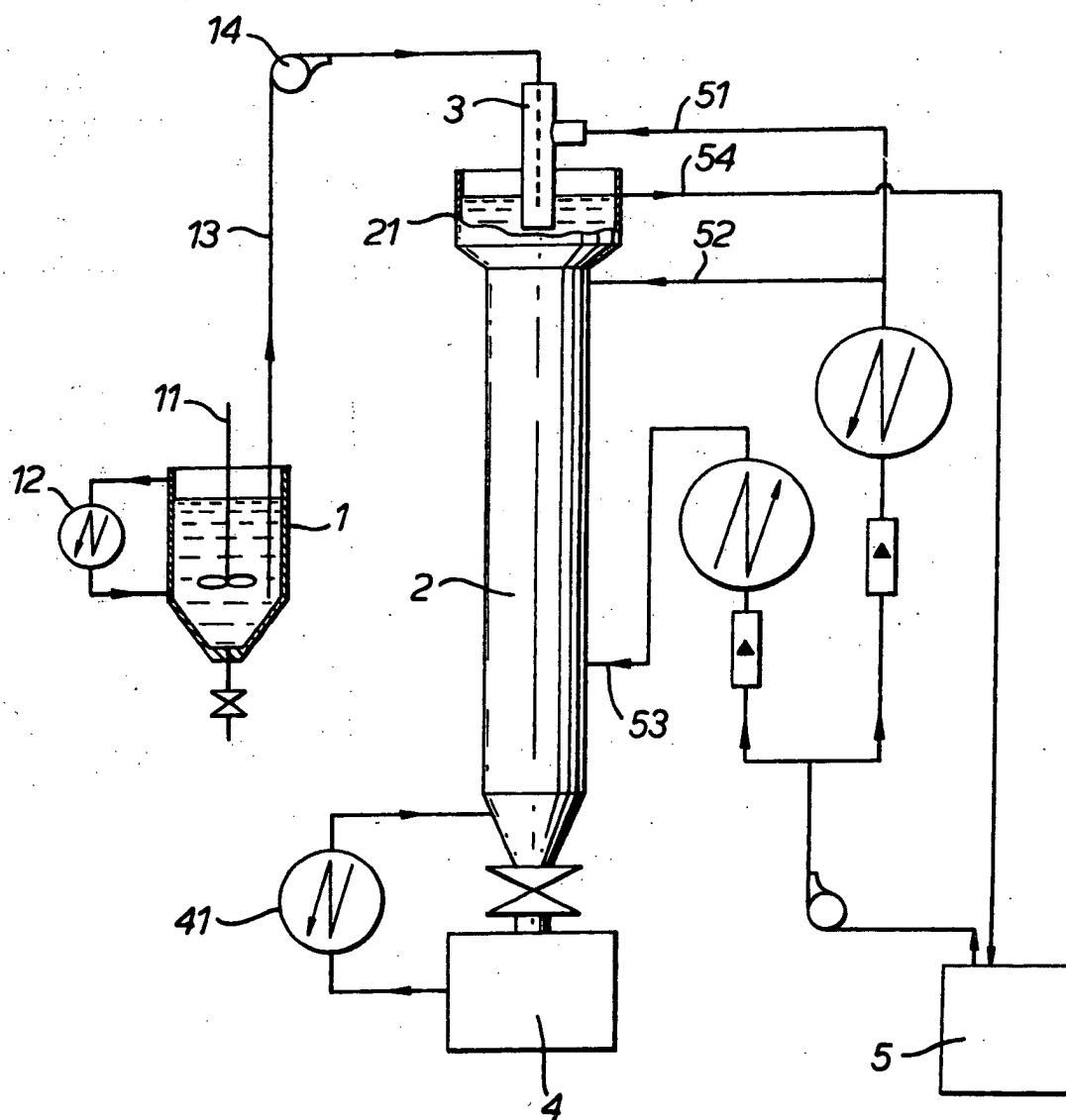
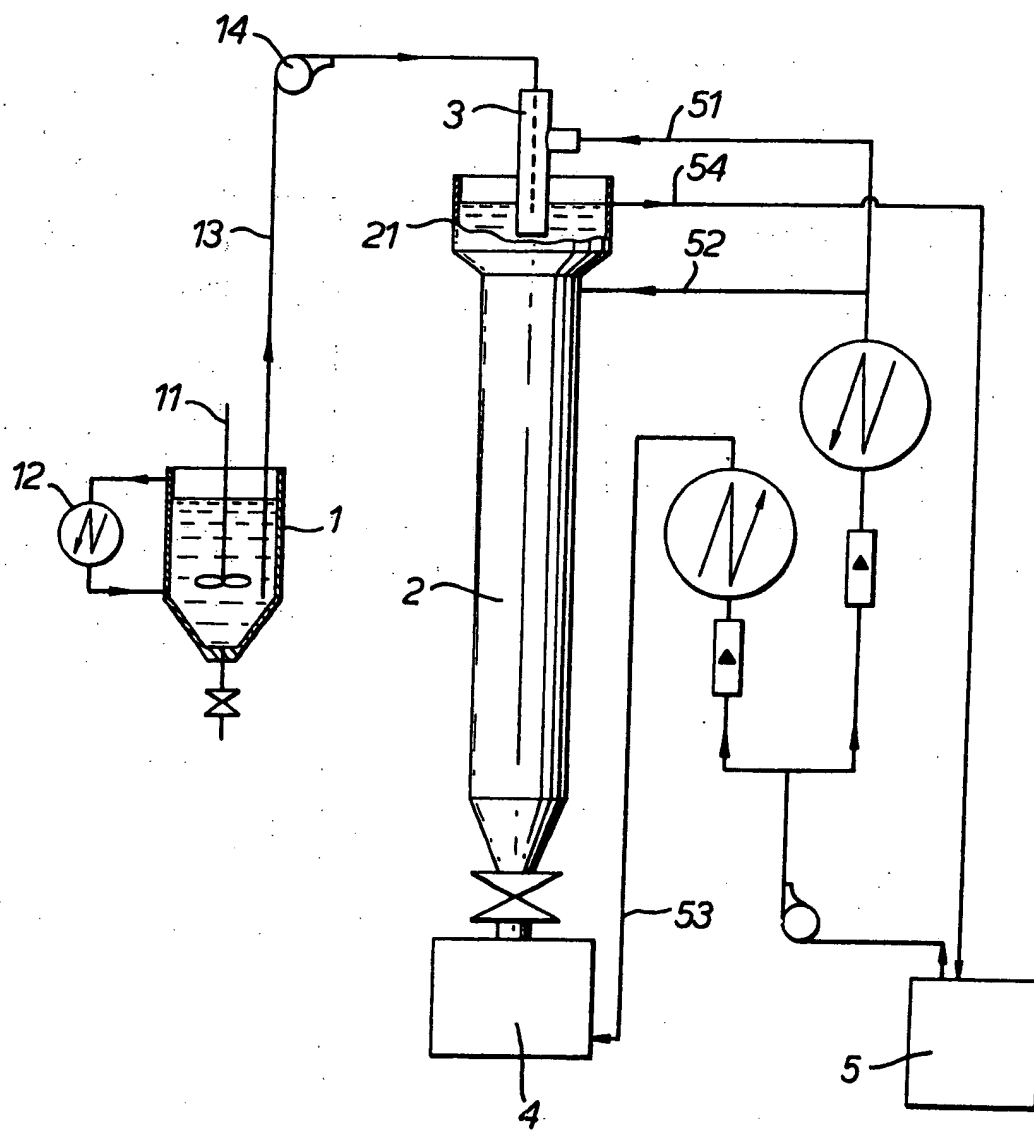


Fig. 2.



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